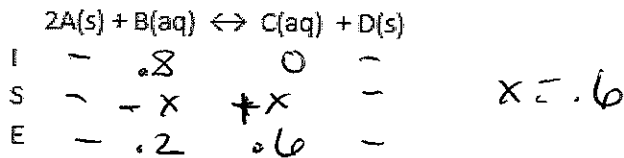


Quiz 4 Topic Reminder
 (#11-2) Can I fill out and calculate ICE tables?



1. In a beaker solid A is placed ~~into a beaker~~ with 0.8M B after a period of time the concentration of B is 0.2 M and appears to be static.

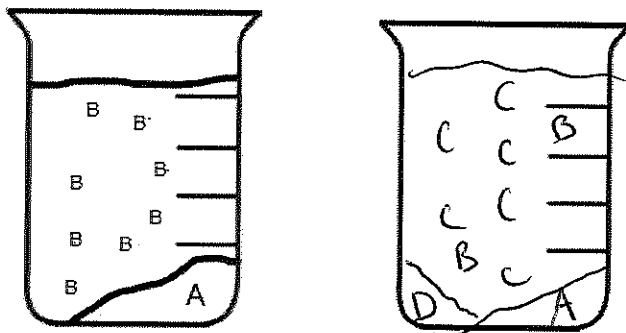
a. Write an equilibrium expression for K_c .

$$K_c = \frac{[C]}{[B]}$$

b. Create an ICE table, solving for each concentration at equilibrium. (above)

$$K = \frac{.6}{.2} = 3$$

d. Draw a proportional picture of the equilibrium particulate drawing.



e. At an alternate temperature the reaction above causing the equilibrium constant $k = 10$. If the initial concentration of B = 0.5M then calculate the following equilibrium concentrations.

B = .05 C = .45

$$10 = \frac{x}{.5-x}$$

$$10(.5-x) = x$$

$$5 - 10x = x$$

$$5 = 11x$$

$$x = \frac{5}{11} \quad x = .45$$

	$2A(s)$	$+ B(aq)$	\rightleftharpoons	$C(aq)$	$+ D(s)$
I	-	.5		0	-
S	-	-x		+x	-
E	-	.5-x		x	-